

# **A COMPARISON OF VARIOUS FACTORS AFFECTING ANALYSIS OF RADON IN WATER BY LIQUID SCINTILLATION COUNTING<sup>1</sup>**

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## **Abstract**

Throughout the United States, laboratories use different sampling methods (Direct Fill vs. Submerged Bottle), sample preparations (Simultaneous Drawing vs. Separate Drawing), scintillators (Mineral Oil vs. Opti-Fluor), and liquid scintillation counting assays (0-2000 keV vs. 130-700 keV) for analyzing radon in water. We compared these variables on the recovery of radon from two “Proficiency Test (PT)” samples and four “Household Well Water” samples. The “130-700 keV” assay had significantly higher radon recovery than the “0-2000 keV” assay. The Direct Fill sampling produced significantly lower results than the “Submerged Bottle” sampling. “Simultaneous Drawing” of both scintillator and water sample yielded higher radon recovery than “Separate Drawing”. “Mineral Oil” scintillator provided higher radon activity than “Opti-Fluor”. However, in eight consecutive measurements of the PT samples at 60 days (full ingrowth) interval, “Mineral Oil” always overestimated the radon activity compared to the predicted/ assigned value, whereas “Opti-Fluor” invariably produced results close to the predicted/assigned value.

## **Introduction**

Both groundwater and surface water in parts of Georgia and South Carolina have long been known to contain various levels of radionuclides from three naturally occurring decay series headed by  $^{238}\text{U}$ ,  $^{232}\text{Th}$ , and  $^{235}\text{U}$  (Cline et al., 1983; Hess et al., 1985; Zapecza and Szabo, 1988; Coker and Olive, 1989). For example, Coker and Olive (1989) tested 90 wells in Georgia for radon ( $^{222}\text{Rn}$ ) and other radionuclides, and concluded that groundwater from the granite and gneiss aquifers in the Piedmont physiographic province contained the highest average concentrations of naturally occurring radionuclides. Likewise, Stone et al. (2002) found elevated levels of radium in drinking water in the “Piedmont and Coastal Plain Sandhills” regions and elevated uranium in water in the “Piedmont (and Blue Ridge)” region of South Carolina.

Furthermore, Albertson (2003) reported activities of gross alpha, radium-226, and combined

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radium-226 plus radium-228 above the Environmental Protection Agency's (EPA) corresponding drinking water Maximum Contaminant Levels (MCL) in community water systems in the Piedmont, Blue Ridge, and parts of the Coastal Plain physiographic provinces of Georgia. Elevated uranium concentrations were also detected in drinking water in the Piedmont and Blue Ridge physiographic provinces (Albertson, 2003).

In 2010, routine testing of private drinking water wells conducted at the University of Georgia's (UGA) Agricultural and Environmental Services Laboratories (AESL) detected uranium at concentrations above EPA's MCL of 30 parts per billion (ppb). As of April 25, 2017, the total number of water samples tested for uranium was 1246. Of these, 150 had detectable amounts of uranium (above 10 ppb) with 64 being above the 30 ppb MCL. One of the wells tested as high as 6297 ppb, which is more than 200 times greater than EPA's MCL. All 64 samples with uranium above the MCL were collected from the Piedmont Blue Ridge Regions above the "Fall Line", which extends from Columbus in the west, Macon in the middle to Augusta in the east. The AESL developed a mapping program and made it available online for public use at: <http://aesl.ces.uga.edu/water/map/>. Sonon et al. (2017) also showed that a high level of uranium in well water was often associated with high levels of radon in well water and in indoor air (measured in air by the homeowner). The geographical association of high uranium and radon results in the AESL database suggests that uranium, radon, and other radionuclides in deep wells originate predominantly from naturally occurring granitic bedrock located primarily in the Piedmont and Blue-Ridge (PBR) regions of Georgia, corroborating the observations of earlier investigators (Coker and Olive, 1989; Stone et al. 2002; Albertson, 2003).

Exposure to radionuclides through drinking water could result in various public health concerns. In this context, alpha radiation from uranium, radium, and their progenies, including radon are particularly important. According to World Health Organization (WHO) when the gross-alpha ( $\alpha$ ) activity in drinking water exceeds 0.5 Bq/L (13.5 pCi/L) or gross-beta ( $\beta$ ) activity exceeds 1 Bq/L (27 pCi/L), radionuclide-specific activities should be analyzed and brought below the WHO guidance levels, which are 0.1 Bq/L for  $^{228}\text{Ra}$ ; 1 Bq/L each for  $^{223-226}\text{Ra}$ ,  $^{234}\text{U}$ , and  $^{235}\text{U}$ ; 10 Bq/L for  $^{238}\text{U}$ ; 100 Bq/L for  $^{222}\text{Rn}$ , and 15  $\mu\text{g/L}$  (ppb) for total uranium (WHO 2004). Ingested radionuclides could get absorbed into the blood stream and accumulate in specific tissues, bone, and organs causing damage to them or can be excreted out of the body (International Commission on Radiological Protection, 2007). Some examples in this regards are: 66% of absorbed uranium could be rapidly eliminated via urine, while the rest could be distributed and stored in the kidney (12–25%), bone (10–15%), and soft tissues (Wrenn et al., 1985); radium could accumulate primarily in the bone (Wrenn et al., 1985). Ingested uranium primarily causes chemical toxicity, especially nephrotoxicity (Zamora et al., 1998; Zamora et al., 2009), whereas ingested radium and radon could induce radiotoxicity leading to cancer (Wrenn et al., 1985). Ingested radon gas through drinking water diffuses into the stomach wall and can cause stomach cancer by irradiating stomach wall tissues (Hopke et al., 2000). In contrast, lung cancer due to inhaled radon from indoor air has been undisputedly established (Darby et al., 2005).

Interestingly, radon in a household water supply poses both inhalation and ingestion risks with most risk due to inhalation of radon released from the water into the indoor air during showering, laundering, etc. A *very rough rule of thumb* for estimating the contribution of radon in household

water to indoor air radon is that water with 10,000 pCi/L of radon contributes about 1 pCi/L to the level of radon in the indoor air. Based on a National Academy of Sciences report on radon in drinking water (NAS, 1999), it has been estimated that in the USA radon in drinking water causes about 168 cancer deaths per year, 89% from lung cancer caused by breathing in radon released from water, and 11% percent from stomach cancer caused by ingesting radon-containing water (USEPA, 2012).

From 2010 through 2013, the UGA Extension conducted a public education program, along with a half-price water testing service, to encourage well owners to test their waters for uranium. The intent was to expand the database to better understand the nature and extent of the problem, and increase public awareness in this regard. The program is still active but at a full price for the water test. People from Georgia counties with high uranium and radon in well water reported (Perry, 2013) to various agencies numerous health problems including cancer, kidney problems, autoimmune disorders, gastrointestinal symptoms, and neuropathy despite the apparent absence of any published report for the affected areas directly linking these contaminants and the reported illnesses. In August 2015, the AESL and the College of Family and Consumer Sciences of the UGA launched a new Radon in Household Water Testing and Education program in collaboration with the Laboratory of Inorganic and Nuclear Chemistry, New York State Department of Health. The EPA recommended Liquid Scintillation Counting (LSC) for analyzing of radon in water (Whittaker et al., 1989), which was subsequently approved in New York State by the Environmental Laboratory Approval Program (NYSDOH, 2007; Kitto et al., 2008a). However, laboratories analyzing radon in drinking water across the USA follow different sampling methods (Direct Fill vs. Submerged Bottle), sample preparations (Simultaneous Drawing vs. Separate Drawing), scintillation fluids (Mineral Oil vs. Opti-Fluor), volume of water plus scintillator in the counting vial (8 mL and 8 mL vs. 10 mL and 10 mL combinations), mixing scintillation fluid and water in the laboratory vs. on-site, and LSC assays (0-2000 keV vs. 130-700 keV). The advantages and disadvantages of these practices are yet to be fully evaluated to optimize the sampling and analysis conditions. Moreover, it has also remained unresolved as to whether presence of an air bubble in the water sample vial is acceptable or unacceptable. This paper reports the results of a study comparing these variables on the recovery of radon from two “Proficiency Test (PT)” samples and four “Household Well Water” samples.

## Methodology

### Water Samples

We used two  $^{222}\text{Rn}$  in water “proficiency test (PT)” or standard samples designated as “Standard-15” and “Standard-17” obtained from co-author Kitto and numerous samples from four household wells from Monroe County, Georgia. The four household wells included in this study are designated as “Well-1: 6090DFR”, “Well-2: 711JP”, “Well-3: 715JP”, and “Well-4: 578TR”. Three of these household well waters contained elevated levels of uranium. A prior analysis of the Well-6090DFR water revealed 629 ppb uranium, 3.8 pCi/L radium ( $^{226}\text{Ra}$  plus  $^{228}\text{Ra}$ ), and 79,000 pCi/L radon. The “Standard-15” and “Standard-17” are reusable radon-in-water standards as they were prepared using a  $^{226}\text{Ra}$ -loaded filter sandwiched in polyethylene sheeting (Kitto et al., 2008b). At full ingrowth (>30 days), the  $^{222}\text{Rn}$  produced by the sandwiched  $^{226}\text{Ra}$  sources in both “Standard-15” and “Standard-17” should be 4375 pCi/L at 100% emanation, but due to retardation by the polyethylene, produces only 3762 pCi/L at 86% emanation. Using these two standard water samples and four well water samples, we compared the seven variables

concerning sampling methods, sample preparations, scintillation fluids, volume of water plus scintillator in the counting vial, and LSC counting assays as described hereunder.

### 1. *Methods of sampling*

We compared “Submerged Bottle” versus “Direct Fill” methods of sample collection. Two samples were collected for each method. For the “Submerged Bottle” method (Figure (1)), water was collected into a bowl by gently flowing down one side with minimal disturbance. Then the entire sample bottle and lid were submerged under water in the bowl, opened, and filled. The water-filled bottle was capped while still under water and turned it upside down. If air bubbles were present, the bottle was emptied and refilled again until air bubbles were no longer observed. The procedure was repeated with the second bottle. For the “Direct Fill” method (Figure (2)), gently flowing water was collected directly into the top opening of the sample bottle, carefully avoiding turbulence. The bottle was allowed to gently overflow, forming a slight dome of water at the opening. The bottle was promptly capped and checked for air bubbles by inverting bottle and tapping gently. If air bubbles were present, the bottle was emptied and the filling procedure repeated until air bubbles were no longer observed in the water sample. The procedure was repeated with the second bottle.



Figure (1): Submerged bottle method of sampling.



Figure (2): Direct-fill method of sampling.

### 2. *Liquid scintillation counting (LSC) assays*

Using a Tricarb 2910 Liquid Scintillation Counter (PerkinElmer, Waltham, MA) for counting radon in water, we compared radon recovery from two different liquid scintillation counting assays as presented in Table 1 below.

Table (1): The two different LSC assays compared in this study.

Regions	Assay-1		Assay-2	
	Lower Limit (keV)	Upper Limit (keV)	Lower Limit (keV)	Upper Limit (keV)
A	0	2000	130	700
B	0	2000	150	1800
C	0	2000	0	2000

The “Assay-1” is a full spectrum assay covering the whole range of energy with the region of interest (ROI) from 0 to 2000 keV. In contrast, the “Assay-2” is limited within the ROI for  $^{222}\text{Rn}$  from 130 to 700 keV, excluding the counts below 130 keV (which is indeed from “Bremsstrahlung” radiation). Cutting out the low-energy (below 130 keV) betas also reduces the

quenching and background. The efficiency (cpm/dpm) in Assay-2 is 3.0 to 3.1 (or about 66% absolute efficiency for each alpha or beta emission).

### 3. *Prepared on-site versus in the laboratory*

Here we compared two situations: (1) right after collection, the samples were injected underneath the preloaded scintillator in the vial “on-site”; and (2) the collected samples were brought to laboratory and then injected underneath the preloaded scintillator in the vial. A 10 mL scintillator and 10 mL sample combination was used in both cases. We compared both “Mineral Oil” and “Opti-Fluor” in combination with these two variables.

### 4. *Type of scintillation fluids*

We compared the efficacy of two different scintillation fluids namely, “Opti-Fluor” and “High Efficiency Mineral Oil Scintillator” (PerkinElmer, Waltham, MA), as both are recommended and used for analyzing radon in water. “Opti-Fluor” used in this study is a benzene-based mixture of high flash point and low volatility organic solvents that produced a background count rate of 15 cpm and 71% quench parameter, it is biodegradable. The “High Efficiency Mineral Oil Scintillator” used in this study contains primarily white mineral oil (60-80%) and 1,2,4-trimethylbenzene (20-40%), it is not biodegradable, with a background of 15 cpm and 107.5% counting efficiency.

### 5. *Volume of sample and scintillation fluid*

We compared recovery of radon for two different preparations: “8 mL scintillation fluid and 8 mL sample” versus “10 mL scintillation fluid and 10 mL sample”. The volumes used by analytical laboratories vary but must be identical to those used to standardize the Liquid Scintillation Counter used for the analyses.

### 6. *Methods of mixing the sample and scintillation fluid*

We compared two different methods of mixing sample and scintillation fluid on the recovery of radon. For the first method, called “Separate Drawing” (Figure (3)), the scintillation fluid (8 or 10 mL) was preloaded into the scintillation vial, and then the sample (8 mL or 10 mL) was pipetted and injected underneath the scintillation fluid. The second method, called “Simultaneous

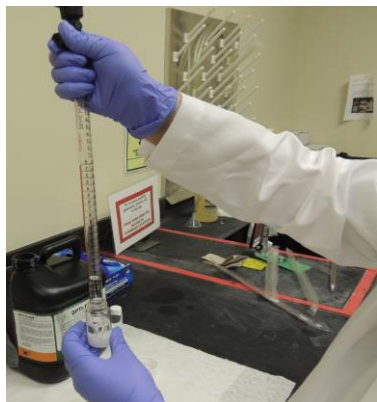


Figure (3): “Separate Drawing” method.

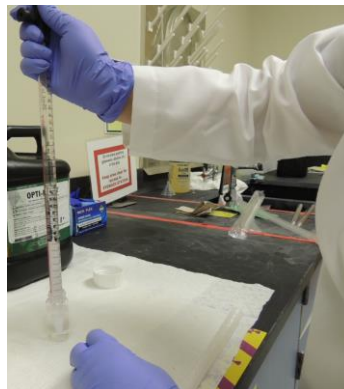


Figure (4): “Simultaneous Drawing” method.

Drawing” (Figure (4)), had the scintillation fluid (8 mL or 10 mL) drawn into a pipette, then the water sample (8 mL or 10 mL) was drawn into the same pipette underneath the scintillation fluid and finally both sample and scintillation fluid was dispensed into the scintillation vial. The vials were capped immediately in both methods and shaken vigorously to expedite transfer of radon into the scintillation fluid.

#### 7. *Effects of air bubble in the sample*

After collection of samples with no air bubble, 0.5 mL of water was quickly withdrawn from a set of samples collected from each of “Well-715JP” and “Well-578TR”. Another set of samples were prepared withdrawing 1 mL of water in the same way. The vials were recapped immediately after such withdrawal. The withdrawal of 0.5 or 1 mL of water created an air bubble of either 0.5 cc or 1 mL thereby enabling us to study the effect of air bubble on radon counts in these two wells. Simultaneous drawing of 10 mL of mineral oil and 10 mL water sample was used to prepare the samples, which were then measured by assay-2 (130-700 keV).

#### Procedure for Testing Uranium in Well Waters

Well water samples were preserved by addition of HNO<sub>3</sub> to pH < 2 upon receipt and filtered prior to analysis when suspended solids appear excessive for passage through the ICP nebulizer. Uranium in the acidified-filtered samples was determined following the EPA Method 200.7 on an “ICP-AVOES” instrument; model “ARCOS FHE” (SPECTRO Analytical Instruments GmbH, Germany). In this method, analysis of samples begins with introduction of the sample into the nebulizer/spray chamber where uniform droplets are swept via an argon gas stream into a high temperature plasma torch. The power delivered by a radio frequency field is absorbed by atomic species in the sample inducing an electronic transition to a higher orbital. Upon passage through this high energy field, the electrons ‘relax’ into more stable orbits by releasing the previously absorbed energy. Much of this released energy is measurable as light in the ultraviolet to visible range of the electromagnetic spectrum. The wavelength of the light emitted corresponds to a specific analyte, and it is 385.958 nm for uranium. The intensity of the light emitted by uranium is positively correlated to concentration in the original sample, and thus the intensity is calibrated to indicate the concentration. It is worth mentioning here that this method is appropriate for the simultaneous analysis of all analytes specified in EPA Method 200.7.

### **Results and Discussion**

It is worth mentioning at the beginning of this section that the bars in all Figures presented and discussed hereafter represent the means (*the main columns*) and standard errors (*the error bars*) from replicated measurements (number of samples included in each case has been cited in each Figures). The standard error (SE) was calculated as:

$$SE = SD \div \sqrt{n} \quad (1)$$

Where, SD = Standard deviation and n = number of samples included.

It also worth noting that the values in the bars followed by same letter(s) do not differ significantly at 5% level of significance by t-test.

#### Studies on the Well Water Sample

##### *1. Sampling method: Direct-Fill Method versus Submerged Bottle Method*

Duplicate water samples were collected from four private well by both “Direct Fill” and “Submerged Bottle” methods. They were prepared in the laboratory by “Simultaneous Drawing”

of 10 mL Opti-Fluor fluid and 10 mL sample, and then counted by “Assay-2 (130-700 keV)”. As depicted in Figure (5), measured radon in “Direct Fill” sampling was significantly lower than that in “Submerged Bottle” sampling for all four wells. Such results suggest that the “Direct Fill” method of sampling is prone to a substantial loss of radon as compared to “Submerged Bottle” method. It is also worth mentioning here that it is very difficult to collect a sample by “Direct Fill” without formation of an air bubble. Generally, it required several attempts to collect a bubble free sample. This method of sampling was dropped from our laboratory’s recommendation because it is troublesome and susceptible to substantial loss of radon during sampling.

2. Liquid scintillation assays: “assay-1 (0-2000 keV)” versus “assay-2 (130-700 keV)”

Figure (6) shows the results of radon concentrations in the water sampled by the submerged bottle method from well-6090DFR given by the two LSC assays, “Assay-1 (0-2000 keV)” versus “Assay-2 (130-700 keV)”.

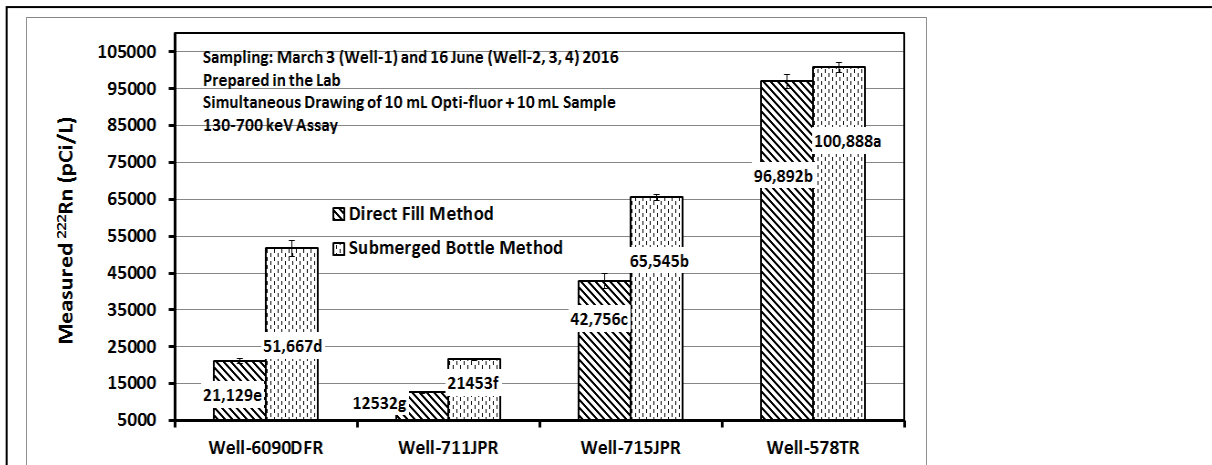


Figure (5): Measured radon in a well water as affected by two different methods of sampling.

- Well-1: 6090DFR; Well-2: 711JPR; Well-3: 715JPR; Well-4: 578TR.

Note: Each bar represents the mean and standard error from duplicate samples.

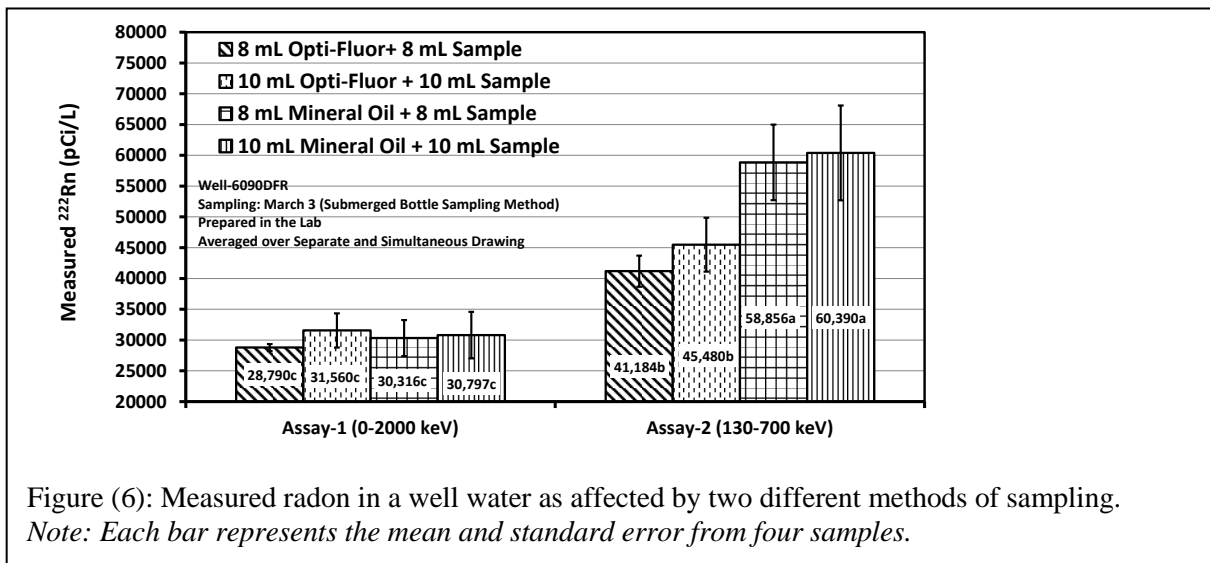


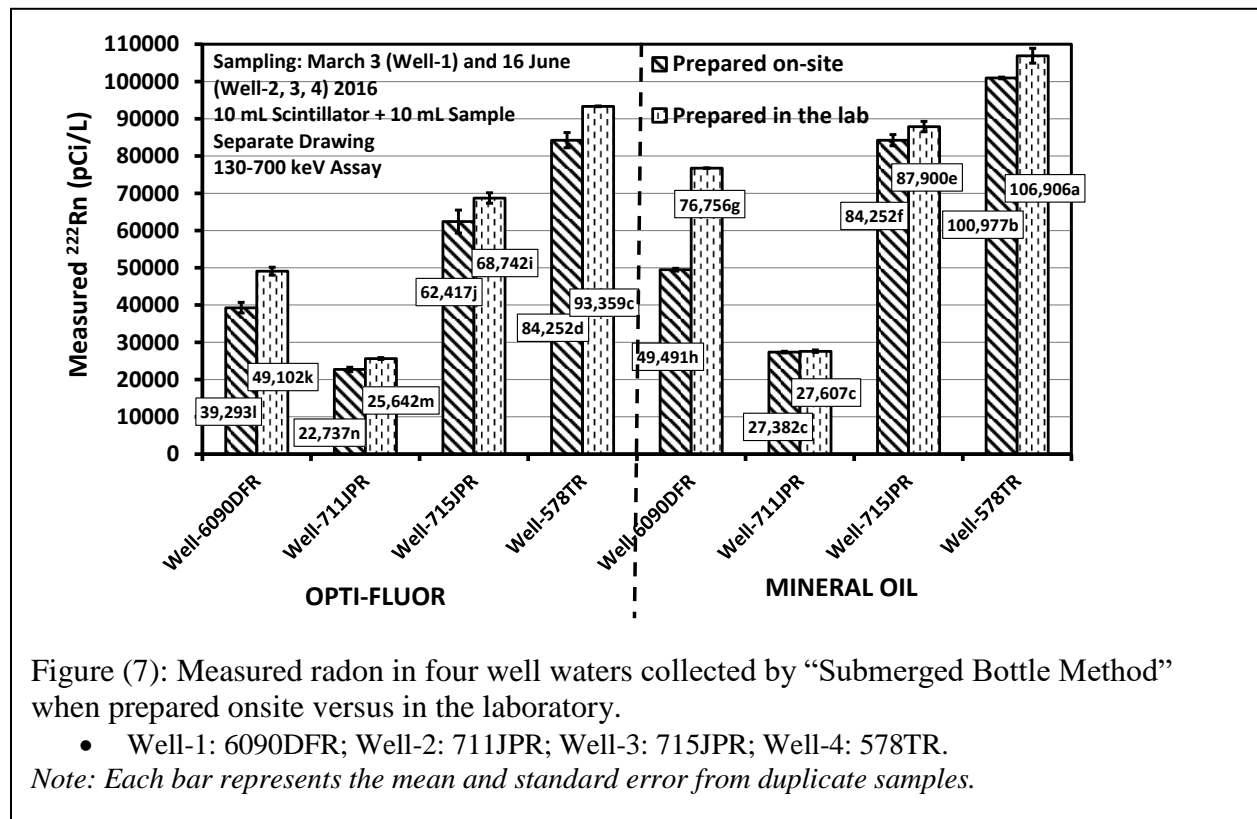
Figure (6): Measured radon in a well water as affected by two different methods of sampling.

Note: Each bar represents the mean and standard error from four samples.

The “Assay-2” gave significantly higher radon results than “Assay-1” regardless of scintillation fluid (Opti-fluor or mineral oil) or volume ratios of the scintillator : sample (10 mL : 10 mL or 8 mL : 8 mL) used. Thus, it is clear that exclusion of the low-energy portion (prior to 130 keV) with higher background and Bremsstrahlung radiation is a better way to analyze radon in water. Similar results were also obtained with the samples from the other three wells (data not shown for the sake of brevity).

### 3. Sample preparation: onsite versus laboratory

Duplicate samples, collected by the “Submerged Bottle” method from each well, were prepared onsite using 10 mL of mineral oil or Opti-fluor by “Separate Drawing”- ( $2 \times 2 = 4$  samples for each of the four wells). Another set of 4 samples for each well was collected by the same method were brought to the laboratory and prepared in the same way using 10 mL of mineral oil or Opti-fluor fluid in duplicate. All of these 8 samples for each of the four wells were counted on LSC by “Assay-2 (130-700 keV)”. The results revealed that the radon concentration in the samples prepared in the laboratory were in general significantly higher than those obtained for the samples prepared on-site (Figure (7)) regardless of whether Opti-fluor or mineral oil was used as the scintillator. The only exception where these two variables produced statistically similar results was the samples from the well-711JPR prepared in mineral oil. Our results suggest that radon is better retained in the scintillation fluids if the waters are injected under the scintillation fluid in the laboratory rather than onsite. Such results may be due the fact that to some extent longer time was elapsed keeping the vial exposed to air in case of on-site preparation as compared to when it was done in the laboratory because preparation in the laboratory was more convenient as compared to onsite. This finding negates the general belief that water sample for testing radon need to be mixed with scintillator on-site. The finding is indeed favorable for practicality. Because the ordinary homeowners would most likely be unable



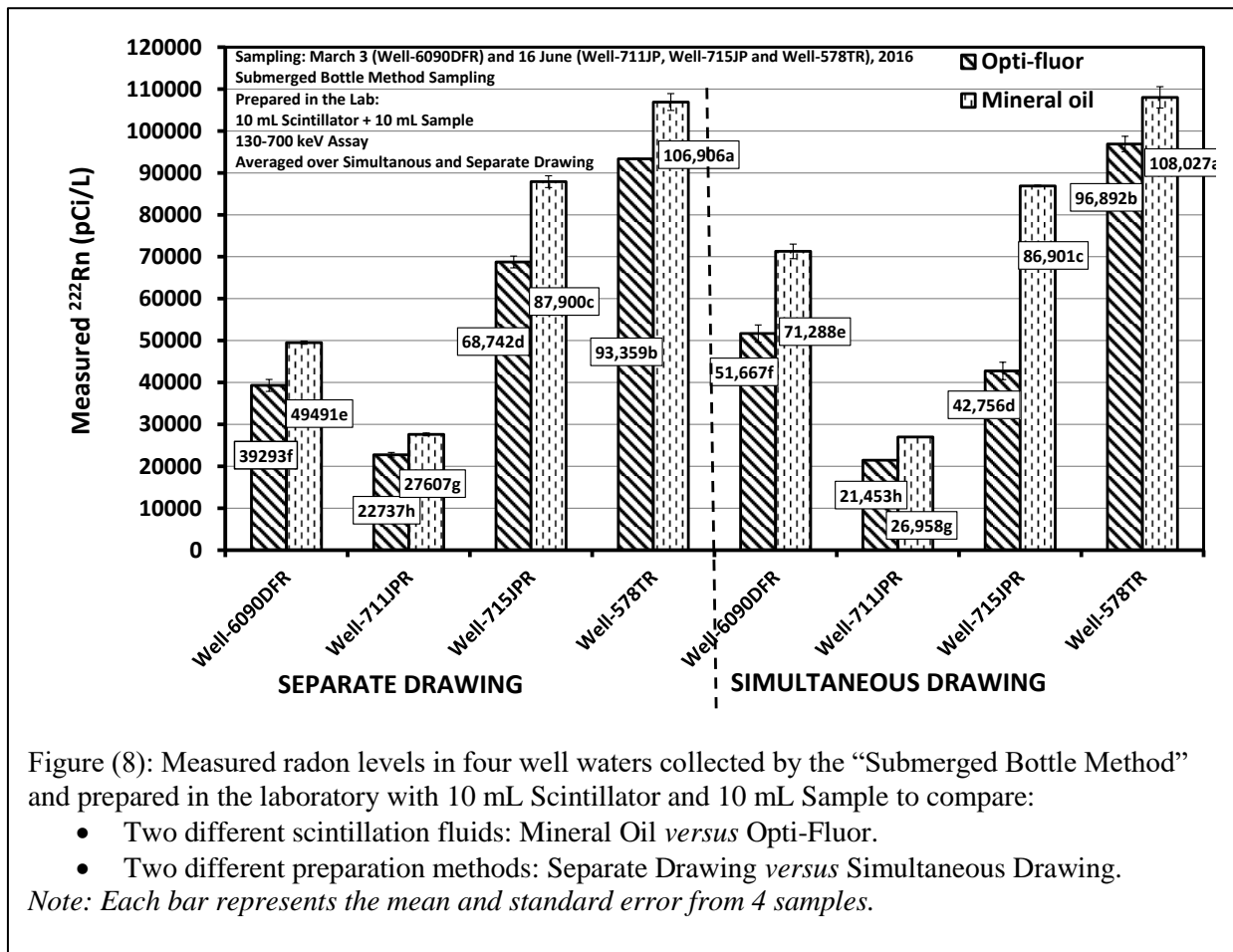


to inject the right amount of water sample under the scintillation fluid in the vial on-site in correct manner by themselves and it would expose them to the scintillation fluid.

4. Effects of scintillation fluid type, volume of scintillation fluid and sample, and methods of mixing

As depicted in Figure (8), for all four well waters, use of 10 mL mineral oil and 10 mL sample resulted in significantly higher radon count rates than 10 mL Opti-fluor and 10 mL sample regardless of the mixing methods (“Simultaneous Drawing” or “Separate Drawing”). Mineral oil also yielded significantly higher radon count rates than Opti-fluor when 8 mL scintillator and 8 mL sample was used (*data not shown*).

The extraction of radon from water relies on the fact that radon is more soluble in organic solvents than in water. When added to water, the Opti-fluor scintillation fluid (emulsifying) initially forms a white emulsion, which should separate into two clear layers before counting to avoid the disruption of extraction and recovery of the radon and interference of water soluble radionuclides, such as radium, with radon counting. Mineral oil is more effective to bring about such desired separation into two distinct layers faster (Figure (9)).



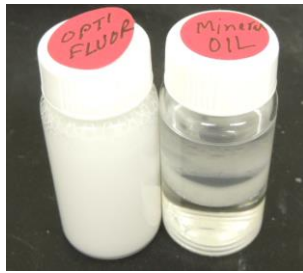


Figure (9a): Sample and Fluid (Opti-Fluor and Mineral Oil) at 3 hours after mixing.



Figure (9b): Sample and Fluid (Opti-Fluor and Mineral Oil) at 5 hours after mixing.



Figure (9c): Sample and Fluid (Opti-Fluor and Mineral Oil) at 68 hours after mixing.

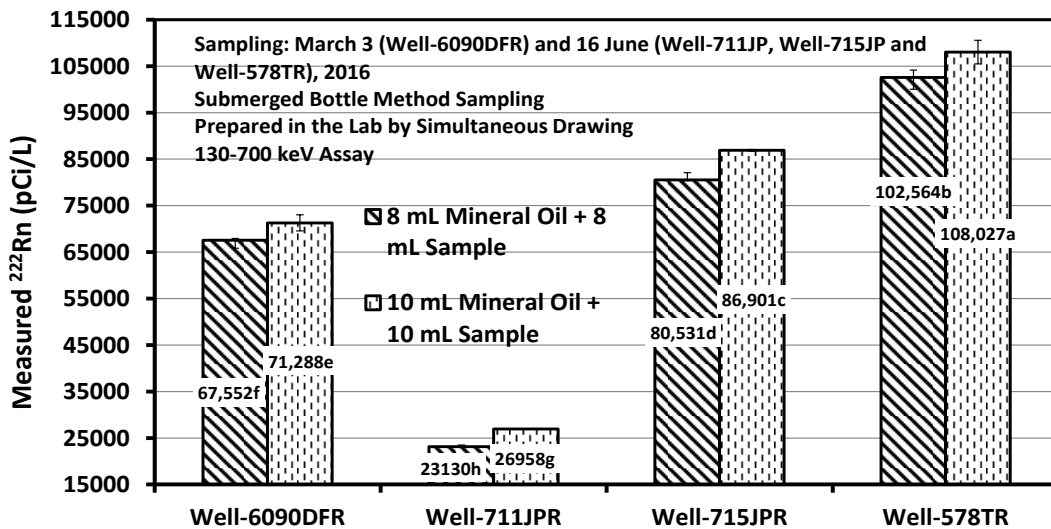


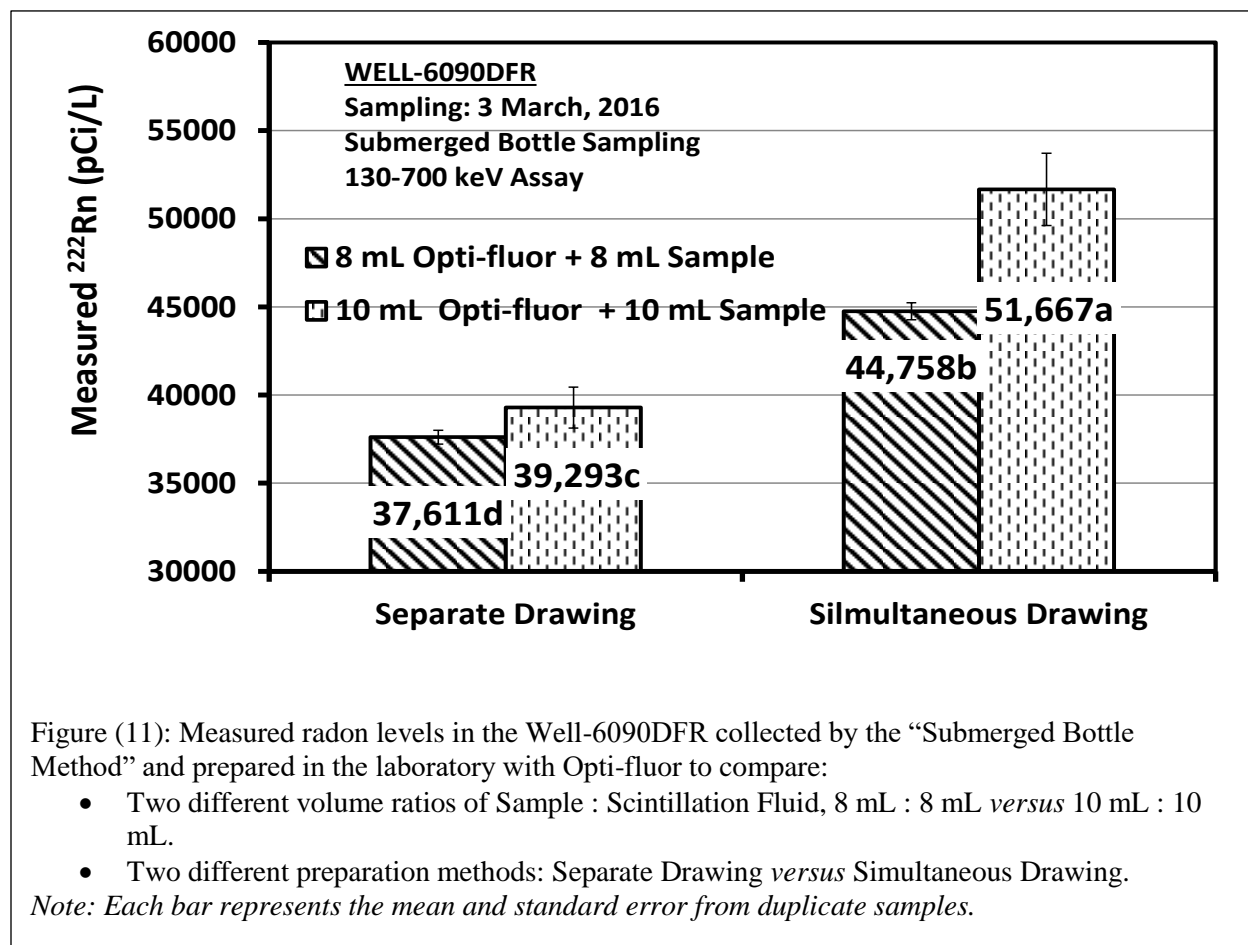
Figure (10): Measured radon levels in four well waters collected by the “Submerged Bottle Method” and prepared in the laboratory by simultaneous drawing of mineral oil and sample to compare:

- Two different volume ratios of Sample: Scintillation Fluid, 8 mL : 8 mL versus 10 mL : 10 mL.

Note: Each bar represents the mean and standard error from duplicate samples.

Figure (10) shows that with simultaneous drawing, 10 mL mineral oil and 10 mL sample resulted in significantly higher radon count rates (i.e., counts per unit volume of water) than 8 mL mineral oil and 8 mL sample for all four well waters. With separate drawing of mineral oil and sample similar results were obtained with regards to the effects of these two different volume ratios (*data not shown*). Furthermore, use of 10 mL Opti-fluor and 10 mL sample also gave significantly higher radon counts rates than 8 mL Opti-fluor and 8 mL regardless of simultaneous or separate drawing (Figure (11)).

The simultaneous drawing of scintillator and the sample in the same pipette yielded significantly higher radon count rates than their separate drawing regardless of whether mineral oil (Figure (8)) or Opti-fluor (Figure (11)) was the scintillator and whether the volume ratios of the scintillator : sample was 10 mL : 10 mL or 8 mL : 8 mL. Such results suggest that some radon escaped, probably into the headspace of the pipette, when done as a “Separate Drawing”.



##### 5. Temporal variation of radon concentration in the same well

When the same three wells were sampled and analyzed on two different dates (28 May, 2016 and 16 June, 2016), we obtained remarkably different radon concentrations for the two different dates in any given wells (Figure (12)). Such results suggest that temporal variation in radon concentration in the well waters is significant and should be duly considered for determining the actual overall exposure of the household members.

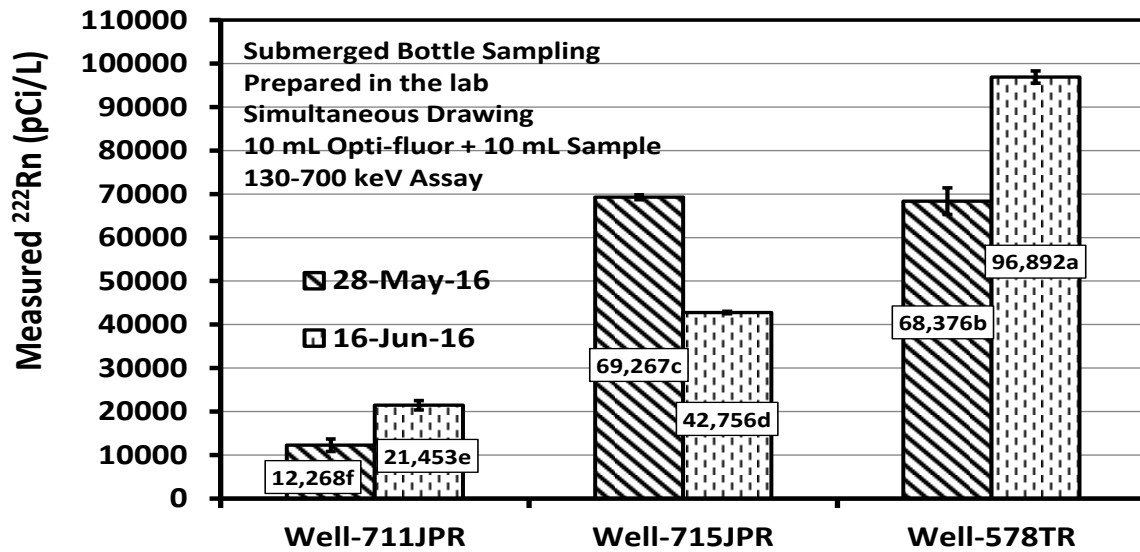


Figure (12): Measured radon levels in three well waters collected by the “Submerged Bottle Method” and prepared in the laboratory by simultaneous drawing of Opti-fluor and sample to compare:

- Two different dates of sampling in a given well.

Note: Each bar represents the mean and standard error from duplicate samples.

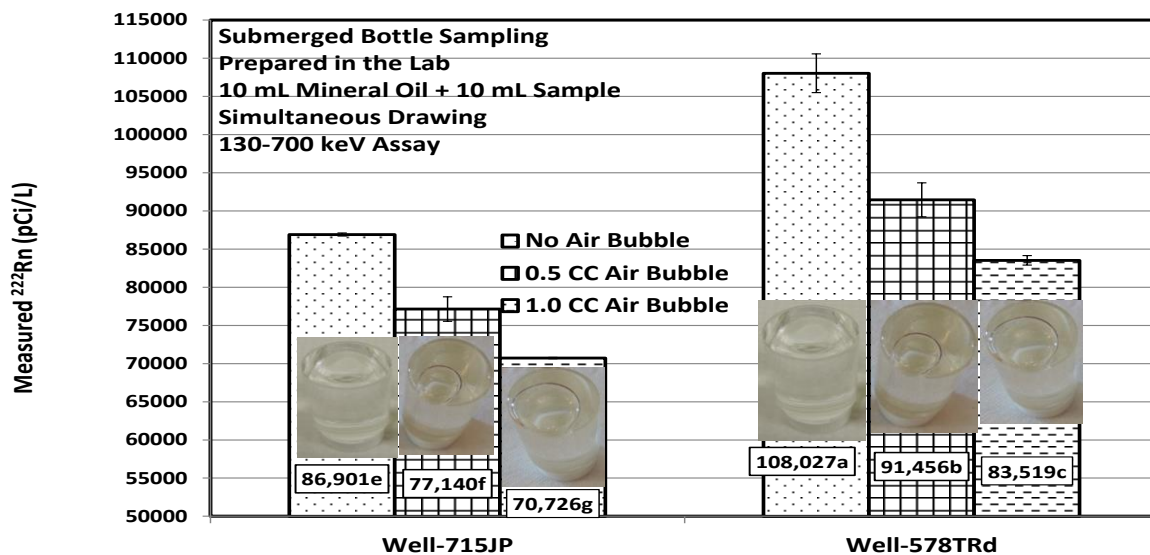


Figure (13): Measured radon levels in two well waters collected by the “Submerged Bottle Method” and prepared in the laboratory by simultaneous drawing of 10 mL mineral oil and 10 mL sample to compare:

- Presence versus absence of air bubble.

Note: Each bar represents the mean and standard error from duplicate samples.

### 6. Effects of air bubble in the sample

The negative effect of air bubble was paramount. As shown in Figure (13), the presence of air bubble gave significantly lower radon counts than when air bubble was absent. This effect was greater as the size of the air bubble was larger. As an example, the well-578TR with 108,027 pCi/L radon, a 1 cm<sup>3</sup> air bubble reduced the radon count by 24,508 pCi/L. It is well known that there could be 4 times as much radon in the air bubble as it would be in an equal volume of water. Thus, for a 22 mL vial used in this study, 1 mL water would contain radon equivalent to 4910 pCi/L for the well-578TR. Therefore, a 1 cm<sup>3</sup> bubble could cause a theoretical loss of 19,640 pCi/L. However, the measured loss (24,508 pCi/L) was to some extent higher than this theoretical loss. Based on this study, a sample with an air bubble must be rejected even though some literature has reported an insignificant effect of an air bubble on the recovery of radon in water.

### Studies on the Standard Samples

The assigned known concentration of radon in both “Standard-15” and “Standard-17” is 3840 pCi/L with the lower and upper acceptance limits of 2880 (~75% of the known) and 4800 (~125% of the known) pCi/L, respectively (Kitto et al, 2008b). As depicted in Figure (14), the assay-1(0-2000 keV) gave unacceptably lower radon counts (2,451-2,532 pCi/L) than that (4,150-5,133 pCi/L) given by assay-2 (130-700 keV) for both Standard-15 and Standard-17 regardless of using mineral oil or Opti-fluor and simultaneous or separate drawing. The results

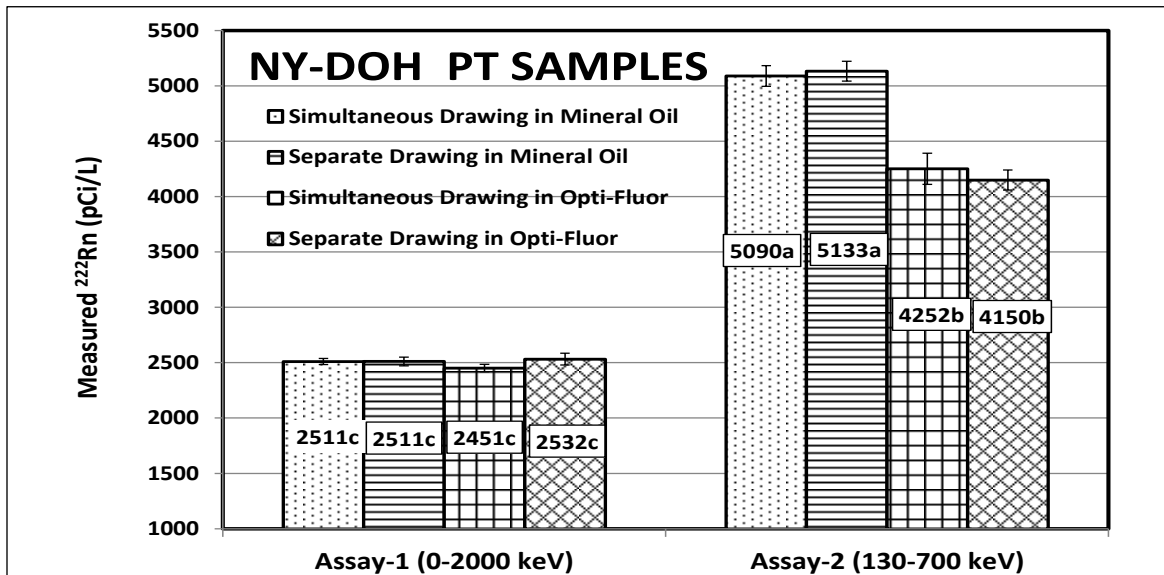


Figure (14): Measured radon levels after 60 d ingrowth in the proficiency testing standard samples received from New York State Department of Health to compare:

- Two different LSC assays: 0-2000 keV versus 130-700 keV.
- Two different scintillation fluids: Mineral Oil *versus* Opti-Fluor.
- Two different preparation methods: Separate Drawing *versus* Simultaneous Drawing.

Note: Each bar represents the mean and standard error from 8 consecutive measurements carried out at 60 days interval.

further show that when measured using assay-2 (130-700 keV), use of “mineral oil” gave higher radon activity than “Opti-fluor” (Figure (14)), just as we observed with the well waters (Figure (8)). However, “simultaneous drawing” gave radon counts statistically identical as “separate drawing” (Figure (14)) for the standard samples despite the fact that “simultaneous drawing” gave significantly higher radon counts than “separate drawing” for the well water samples (Figure (8)). Such variable results may be related to the fact that the radon concentrations in the four well waters (21,453-108,027 pCi/L) were much higher than that in the standard samples (4,150-5,133 pCi/L). However, it worth noting that “mineral oil” grossly overestimated the radon activity (5090 and 5133 pCi/L) as compared to their predicted or assigned value (3840 pCi/L).

Figure (15) plots the detailed results of the 8 consecutive radon measurements carried out at 60 day intervals at the University of Georgia (UGA) laboratory using LSC assays and compare these results with those reported by other laboratories across the United States using various methods. The x-axis represents the various laboratories in the US. The y-axis is the other laboratories in the nation (Kitto, 2008b) and plotted in ascending order. The data points corresponding to x-axis values 23-94 are the values obtained in the UGA laboratory by 4 different sample processing (mineral oil and Opti-fluor in combination with simultaneous and separate drawings:  $2 \times 2 = 4$ ). The blue-filled circles are the results obtained from the assay-2 (130-700 keV ROI), whereas the red-filled circles are the results obtained from the assay-1 (0-200 keV ROI).

As depicted in Figure (15), all results generated by the assay-1 (0-2000 keV ROI) were lower than the lower limit of acceptance, which means they all failed. Among the results generated by assay-2 (130-700 keV), the data points for the samples prepared in “mineral oil” were mostly higher than the upper limit of acceptance, which means they also failed in most cases. In sharp contrast, all results from Opti-fluor are acceptable, and at least half of them were very close to the true value. Therefore, the assay-1 (0-2000 keV ROI) with both mineral oil and Opti-fluor can grossly underestimate the actual radon concentration, and with assay-2 (130-700 keV ROI), mineral oil can over-estimate the radon concentration, thus both of these should be avoided.

#### Uranium concentration

Out of the four wells, three wells (Well-6090DFR, Well-715JP, and Well-578TR) had uranium concentration much higher than the EPA’s drinking water MCL (Table 2). The uranium level in the Well-711JP was below the MCL. However, all of the four study wells had radon concentrations much higher than the EPA-proposed MCL of 300 pCi/L as well as the proposed alternate MCL (AMCL) of 4000 pCi/L. Furthermore, the radon concentrations followed the same trend as the uranium concentrations in the wells. Nevertheless, it is important to note here that despite there was no concern about the uranium concentration (22.6 ppb) in the well-711JP, it contained well over 20,000 pCi/L radon. Therefore, using uranium concentration over 30 ppb as a trigger for recommending test for radon in well water remains questionable. We also noticed temporal variation in the uranium concentration in the well-6090DFR and well-578TR. Such variation was significant in the case of well-578TR (1,549 ppb in June, 2001 versus 4,939 ppb in June, 2016)

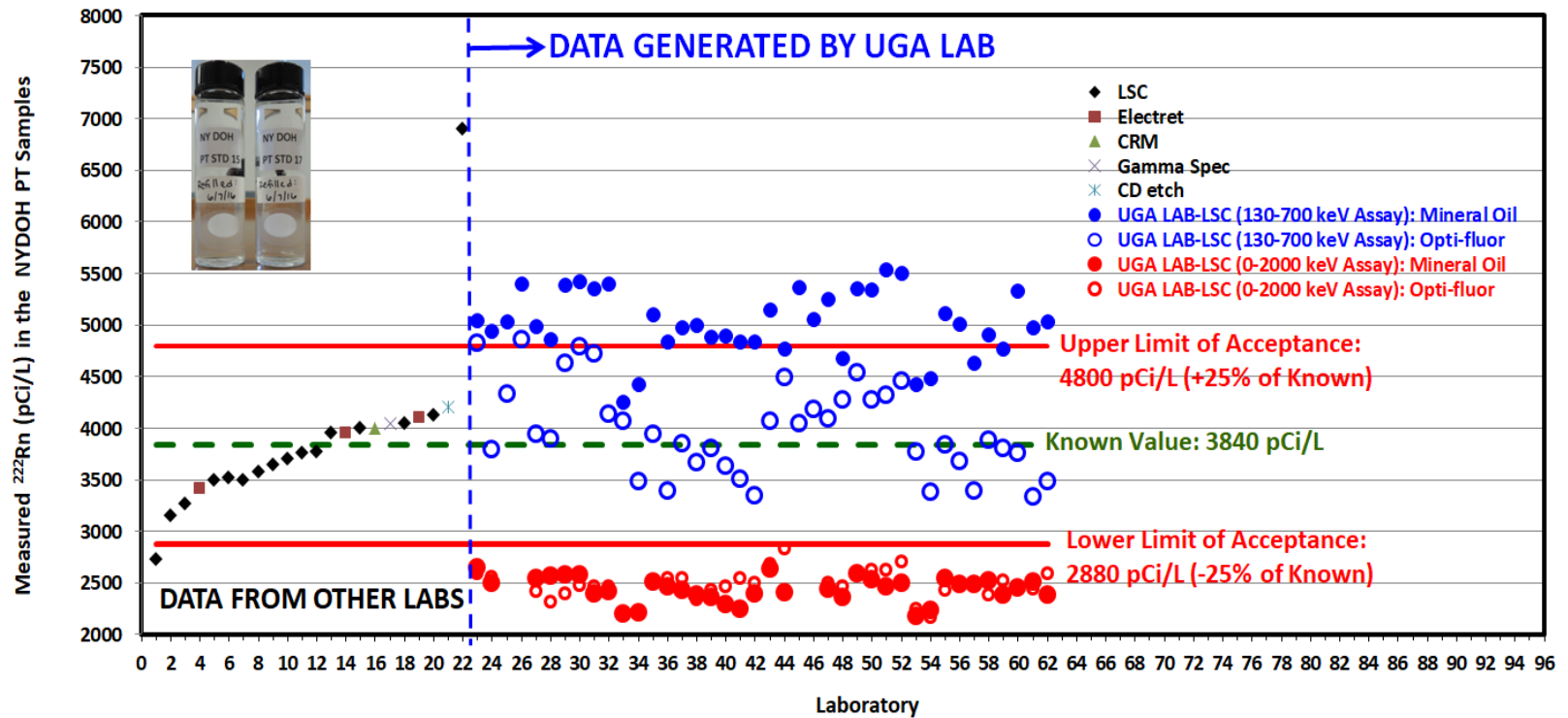


Figure (15): Radon levels in the NYDOH proficiency testing standard samples in 8 consecutive measurements at University of Georgia laboratory to compare:

- Two different LSC assays: 0-2000 keV versus 130-700 keV.
- Two different scintillation fluids: mineral oil *versus* opti-fluor.
- UGA laboratory results by LSC with the other results from different laboratories across the United States generated by various methods.

Table (2). Concentration of uranium in the four wells.

Well	Date	Uranium Conc. (ppb) <sup>‡</sup>	Sampler	Date	Uranium Conc. (ppb)
Well-6090DFR	3 Mar 2016 <sup>†</sup>	542	U. Saha	14 June, 2011 <sup>§</sup>	629
Well-711JP	20 Jun 2016 <sup>†</sup>	22.6	D. Lynch		
Well-715JP	16 June 2016 <sup>†</sup>	407	U. Saha		
Well-578TR	16 Jun 2016 <sup>†</sup>	4,939	U. Saha	15-June, 2011 <sup>§</sup>	1,549

<sup>†</sup>Sampled and analyzed as part of this study.

<sup>‡</sup>EPA's MCL for drinking water is 30 ppb.

<sup>§</sup>Voluntary submission by the well owner.

### Conclusions

- Direct-fill method is susceptible to significant loss of radon gas; so the submerged bottle method should be used.
- The assay-2 (130-700 keV) based on the region of interest for radon is better than the assay-2 based on a full spectrum (0-2000 keV) analysis.
- An air bubble in a sample results in significant loss of radon gas, and such loss becomes greater as the size of the air bubbles becomes larger.
- Mineral oil generally gives higher radon count rates than Opti-fluor. But the results of the two proficiency or standard samples showed that mineral oil clearly overestimates the actual radon concentration whereas Opti-fluor always gave the results close to the predicted or assigned value. As a scintillator for radon in water, it is widely believed that mineral oil is a better scintillator for radon than Opti-fluor. But our results show that the opposite is indeed true.
- A 10 mL scintillator and 10 sample combination is better than 8 mL scintillator and 8 mL sample combination for a liquid scintillation cocktail.
- Separate drawing of scintillator and sample may result in significant loss of radon, so simultaneous drawing should be adopted when laboratories use open pipettes. However, this effect may be insignificant if a closed-top (no headspace) sampling syringe is used.
- Mixing scintillator and sample in the laboratory is better than doing making the cocktail on-site for both better results and practicality of testing radon in private well waters.
- There were noticeable temporal variations in both radon and uranium concentrations in the study wells.
- Use of uranium concentration over 30 ppb (the MCL of uranium in drinking water) as a trigger for recommending test for radon in well water remains questionable because there may be safe level of uranium but unsafe level of radon, and *vice versa* in a given well water.



## References

- Albertson, P.N. 2003. Naturally occurring radionuclides in Georgia water supplies: implications for community water system. In Proceedings of the 2003 Georgia Water Resources Conference, held April 23–24, 2003, at the University of Georgia. Kathryn J. Hatcher, editor, Institute of Ecology, The University of Georgia, Athens, Georgia.
- Cline, W., Adamovitz, S., Blackman, C., Kahn, B. 1983. Radium and uranium concentrations in Georgia community water systems. *Health Physics* 44, 1-12.
- Coker, G., Olive, R. 1989. Radionuclide concentrations from selected aquifers in Georgia. U.S. Environmental Protection Agency Region IV Report, August 1989, 21 pp.
- Darby, S., Hill, D., Auvinen, A., Barros-Dios, J.M., Baysson, H., Bochicchio, F., Deo, H., Falk, R., Forastiere, F., Hakama, M., Heid, I., Kreienbrock, L., Kreuzer, M., Lagarde, F., Mkelinen, I, Muirhead, C., Oberaigner, W., Pershagen, G., Ruano-Ravina, A., Ruosteenoja, E., Rosario, A.S., Tirmarche, M., Tomsek, L., Whitley, E., Wichmann, H.E., Doll, R., 2005. Radon in Homes and Risk of Lung Cancer: Collaborative Analysis of Individual Data from 13 European Case-control Studies. *British Medical Journal* 330(7485), 223–226.
- Hess, C.T., Weiffenbach, C.V., Norton, S.A. 1983. Environmental radon and cancer correlations in Maine. *Health Physics* 45, 339–348.
- Hopke, P.K., Borak, T.B., Doull, J., Cleaver, J.E., Eckerman, K.F., Gundersen, C.S. 2000. Health risks due to radon in drinking water. *Environmental Science and Technology* 34, 921–926.
- International Commission on Radiological Protection. 2007. Publication 100: Human Alimentary Tract Model for Radiological Protection. Oxford: Elsevier Sciences.
- Kitto, M.E., Fielman, E.M., Haines, D.K., Menia, A., Bari, A. 2008a. Performance of a commercial radon-in-water measurement kit. *Journal of Environmental Radioactivity* 99, 1255–1257.
- Kitto, M.E., Menia, T.A., Bari, A., Fielman, E.M., Haines, D.K. 2008b. Development and intercomparison of a reusable radon-in-water standard. *Radiation Measurements* 45, 231–233.
- National Academy of Science (NAS). 1999. National Academy of Science. Report of the Committee on Risk Assessment of Exposure to Radon in Drinking Water. Board on Radiation Effects Research, Commission on Life Sciences, National Research Council, National Academy Press, Washington, DC.
- New York State Department of Health (NYSDOH). 2007. Environmental Laboratory Approval Program. <http://www.wadsworth.org/labcert/elap/radon.html>.
- Perry, J. 2013. HEALTH CONSULTATION – June 30, 2013: Naturally-occurring Uranium in Private Well Water, Juliette, Monroe County, Georgia. Chemical Hazards Program, Georgia Department of Public Health, Atlanta, GA. [https://dph.georgia.gov/sites/dph.georgia.gov/files/EnvHealth/Chemical/Monroe%20County%20HC%206\\_30\\_13.pdf](https://dph.georgia.gov/sites/dph.georgia.gov/files/EnvHealth/Chemical/Monroe%20County%20HC%206_30_13.pdf)
- Stone, P.A., Devlin, R.J., Allen B.P., Crawford, B.D. 2002. Radionuclides in South Carolina well water. Proceedings of the tenth annual David S. Snipes/Clemson Hydrogeology Symposium, April 18, 2002, p. 32.
- Sonon, L., Saha, U, Turner, P., Lynch, D., and Dean, G. 2017. Testing, Mapping, and Public Education on Uranium and Radon in Private Drinking Water Wells in Georgia. In Proceedings of the 2017 ASA-CSSA-SSSA International Annual Meeting, held October 22-25, 2017, at Tampa FL. <https://scisoc.confex.com/crops/2017am/webprogram/Paper106211.html>

- United States Environmental Protection Agency (USEPA).1994. Method 200.7 Determination of metals and trace elements in water and wastes by inductively coupled plasma-atomic emission spectrometry. Rev. 4.4. U.S. Environmental Protection Agency, Cincinnati, Ohio.
- United States Environmental Protection Agency (USEPA), 2012. A Citizen's Guide to Radon, The Guide to Protecting, Yourself and Your Family from Radon. EPA402/K-12/002|May2012|www.epa.gov/radon. Retrieved 2016-05-24:  
[https://www.epa.gov/sites/production/files/2016-002/documents/2012\\_a\\_citizens\\_guide\\_to\\_radon.pdf](https://www.epa.gov/sites/production/files/2016-002/documents/2012_a_citizens_guide_to_radon.pdf).
- Whittaker, E.L., Akridge, J.D., Giovano, J. 1989. Two Test Procedures for Radon in Drinking Water: Interlaboratory Collaborative Study. U.S. EPA Environmental Monitoring Systems Laboratory, Las Vegas, NV. EPA 600/2-87/082.
- World Health Organization (WHO). 2004. Guidelines For Drinking Water Quality. 3rd edition. World Health Organization, Geneva.
- Wrenn, M.E., Durbin P.W., Howard, B., Lipsztein, J., Rundo, J., Still, E.T. 1985. Metabolism of ingested U and Ra. Health Physics 48, 601–633.
- Zamora, M.L., Tracy, B.L., Zielinski, J.M., Meyerhof, D.P., Moss, M.A. 1998. Chronic ingestion of uranium in drinking water: a study of kidney bioeffects in humans. Toxicological Science 43, 68–77.
- Zamora, M.L., Zielinski, J.M., Moodie, G.B., Falcomer, R.A., Hunt, W.C., Capello, K. 2009. Uranium in drinking water: renal effects of long-term ingestion by an aboriginal community. Archives of Environmental & Occupational Health 64, 228–241.
- Zapeczka, O.S., Szabo, Z. 1988. Natural Radioactivity in Ground Water - A Review, pp. 50 - 57 in National Water Summary 1986 - Hydrologic Events and Ground-Water Quality, Water-Supply Paper 2325, U.S. Geological Survey, Reston, VA, 1988.